

Write the Formula / Formula Unit for the following Compounds

Determining the formula for Magnesium Fluoride?

1. Identify the charges = $\text{Mg}^{2+} \text{F}^{-1}$
2. Cross the Charges, $\text{Mg}^{2+} \text{F}^{-1} = \text{Mg}_1\text{F}_2$
3. If the subscript is a 1 it does not need to be written.
4. If there is a common subscript such as 2 as in Mg_2O_2 , reduce it to Mg_1O_1 which is also MgO .

Write Formula Unit For the Below Ionic Compounds

	Name	Cation (+)	Anion (-)	Formula
1	Sodium Chloride	Na^{1+}	Cl^{1-}	$\text{Na}^{1+} \text{Cl}^{1-} = \text{NaCl}$
2	Aluminum Chloride	Al^{3+}	Cl^{1-}	
3	Aluminum Phosphide			
4	Magnesium Oxide			
5	Cesium Fluoride			
6	Strontium Nitride			
7	Lithium Sulfide			
8	Calcium Chloride			
9	Sodium Bromide			
10	Beryllium Iodide			
11	Strontium Fluoride			
12	Aluminum Fluoride			
13	Potassium Nitride			
14	Sodium Sulfide			
15	Lithium Oxide			
16	Calcium Oxide			