Review sheet for Chemical Reactions and Balancing Equations Test

1. Be able to list some evidence that would let you know that a chemical reaction has occurred (Think about the labs we have done – what did you see? What happened?)
2. Know the basic parts of a chemical formula:

Reactants 🡪 Products

The arrow means “yields” or “produces” or “forms”. Recognize that bonds must be broken and new ones formed

1. Be able to classify chemical reactions as: Synthesis, Decomposition, Single displacement, Double displacement or Combustion
2. Combustion reactions require a little activation energy to get them going but they are exothermic reactions. Energy in the form of heat and light is released. What are the products of every combustion reaction? What do you need to burn the hydrocarbon?
3. The Law of Conservation of Matter and the Law of Conservation of Mass requires that chemical equations be balanced.
4. Make sure you can use coefficients to balance chemical reactions.
5. Know the difference between coefficients and subscripts
6. Be able to explain 3-5 ways a chemist can increase the rate of a chemical reaction
7. Recognize that if you don’t have enough of a reactant, you will get limited (or no) product forming. The reagent (or reactant) that affects how much product you get (your yield) is the limiting reagent.