Skill Practice 44

Percent Yield

Practice

Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Date: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Hour: \_\_\_\_\_

1. In a certain chemical reaction, 297 g of zinc chloride was produced from the single replacement reaction of excess zinc and 202.7 g of lithium chloride. What is the percent yield of zinc chloride?
2. When 24.5 g of CaCl2 reacted with plenty of AgNO3, 21.5 grams of Ca(NO3)2 were produced. What was the percent yield of Ca(NO3)2?
3. When 312.7 g of Al2(SO4)3 react with plenty of magnesium, what is the percent yield when 38.5 g of aluminum are formed?
4. Consider the reaction in which 370 g of Ca(NO3)2 react with just the right amount of lithium metal in a single replacement reaction. If the percent yield of LiNO3 is 74.3%, what mass of LiNO3 will be produced?
5. If 42.4g of Al react with excess O2 and 67.3g of Al2O3 are produced, what was the percent yield of Al2O3? Note: the balanced equation is: 4 Al + 3 O2 🡪 2 Al2O3