Skill Practice 43

Limiting Reactants

Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Date: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Practice

Hour: \_\_\_\_\_

1. Consider the following combustion reaction: 2 C4H10 + 13 O2 🡪 8 CO2 + 10 H2O. 125 g of C4H10 react with 415 g of O2.
2. Which substance is the limiting reactant?
3. What mass of CO2 and H2O can be produced?
4. When 412.5 g of calcium carbonate react with 521.9 g of aluminum fluoride, how many grams of each product can be produced?
5. When 277 g of Na2SO4 reacts with 137 g of Ca3P2 according to the following balanced equation… 3 Na2SO4 + Ca3P2 🡪 2 Na3P + 3 CaSO4
6. How many grams of the excess reactant are left over?
7. How many grams of Na3P are produced?
8. If 312 g of potassium sulfide reacts with 410 g of aluminum phosphide, how many grams of each product can be produced?