Skill Practice 41

Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Date: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Hydrates

Hour: \_\_\_\_\_

1. Write the formula for the hydrate of CaSO4 that has a mass of 97.85 g before heating and 64.01 g after heating.
2. 140.4 g of hydrated compound of MgSO4 was heated until it had a mass of 68.6 g. What is the formula of the hydrate?
3. A hydrate of MgSO4 was heated intensely so that the mass changed from 2.389 g to 0.855 g. What is the formula of the hydrate?
4. If 9.99 g of a hydrated CaCO3 are heated so that the mass then becomes 5.81 g, which of the following is the formula of the hydrate?