

Practice Quiz on Chapter 4.

1. Complete the following table.

Symbol	Atomic Number	Mass Number	Protons	Neutrons	Electrons
Na	_____	_____	11	13	_____
_____	15	_____	_____	17	_____
_____	_____	60	_____	_____	27
_____	_____	_____	_____	79	53

2. Give the number of protons, neutrons and electrons for each atom listed in the following table:

Isotope	Protons	Neutrons	Electrons
$^{27}_{13}\text{Al}$			
$^{56}_{26}\text{Fe}$			
$^{30}_{14}\text{Si}$			
$^{142}_{56}\text{Ba}$			

3. Magnesium has naturally occurring isotopes with the isotopic masses and relative abundance.

Isotope	mass (amu)	Relative abundance (%)
Magnesium-24	23.985	78.70
Magnesium-25	24.986	10.13
Magnesium-26	25.983	11.17

What is the average atomic mass of magnesium? Calculate from the data. (show your work).